



Quality Checkers
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XI-SCI : Chemistry
Modern Periodic Table,

DATE:

TIME: 1 hour 30
minutes

MARKS: 25

SEAT NO:

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Note:-

1. All Questions are compulsory.
2. Numbers on the right indicate full marks.

Section A

Q.1 Select and write the correct answer.

(4)

1. Which of the following pairs is NOT isoelectronic?
A) Na^{\oplus} and Na B) $\text{Mg}^{2\oplus}$ and Ne
C) $\text{Al}^{3\oplus}$ and $\text{B}^{3\oplus}$ D) P^{3-} and N^{3-}
2. The first attempt to classify elements was made by
A) Mendeleev B) Newlands
C) Lothar Meyer D) Dobereiner
3. Which of the following has highest ionisation enthalpy?
A) B B) N
C) C D) O
4. In the modern periodic table, the elements are arranged in
A) Increasing mass B) Increasing atomic volume
C) Increasing atomic number D) Alphabetical order

Q.2 Answer the following.

(3)

1. Give some examples of periodic events.
2. Predict the position of the element in the periodic table satisfying the electronic configuration $(n - 1)d^1ns^2$ for $n = 4$
3. How many days pass between two successive full moon nights?

Section B

Attempt any Four

Q.3 The electronic configuration of some elements are given below :

(2)

(a) $1s^2$ (b) $1s^2 2s^2 2p^6$

In which group and period of the periodic table they are placed?

Q.4 Give the main characteristics of modern periodic table.

(2)

Q.5 Consider the oxides Li_2O , CO_2 , B_2O_3

(2)

- (a) Which oxide would you expect to be the most basic?
- (b) Which oxide would be the most acidic?
- (c) Give the formula of an amphoteric oxide.

Q.6 Give reason : Alkali metals have low ionization energies.

(2)

